1. Explore building an atom. Click on the first tab “ATOM”
2. What happens to the element name each time a proton is added?
3. Observe the Net Charge. What causes the atom to be positive, negative or neutral?
4. Note the Mass Number. What are 2 scenarios that make the mass number increase? What particle has no effect on the mass number?
5. Click on the 2nd tab, “SYMBOL”
6. Observe the changes on the symbol square.
7. Which number do the protons and neutrons affect?
8. What must happen to get a positive charge?
9. What must happen to get a negative charge?
10. What do you think causes the stable or unstable atom? What is a possible outcome?
11. Note the periodic table.
12. Observe the changes in the table as protons, neutrons and electrons are added or subtracted.
13. Is there a pattern to the arrangement of the table?
14. Note where the Atomic number, Mass Number, Symbol and Net charge are displayed in the periodic table square.
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